300202 - Q - Chemistry

Coordinating unit: 300 - EETAC - Castelldefels School of Telecommunications and Aerospace Engineering
Teaching unit: 745 - EAB - Department of Agri-Food Engineering and Biotechnology
Academic year: 2018
Degree: BACHELOR'S DEGREE IN AEROSPACE SYSTEMS ENGINEERING (Syllabus 2015). (Teaching unit Compulsory)
BACHELOR'S DEGREE IN AEROSPACE SYSTEMS ENGINEERING/BACHELOR'S DEGREE IN TELECOMMUNICATIONS SYSTEMS ENGINEERING - NETWORK ENGINEERING (AGRUPACIÓ DE SIMULTANEITAT) (Syllabus 2015). (Teaching unit Compulsory)
BACHELOR'S DEGREE IN AEROSPACE SYSTEMS ENGINEERING/BACHELOR'S DEGREE IN TELECOMMUNICATIONS SYSTEMS ENGINEERING (Syllabus 2015). (Teaching unit Compulsory)
BACHELOR'S DEGREE IN AEROSPACE SYSTEMS ENGINEERING/BACHELOR'S DEGREE IN NETWORK ENGINEERING (Syllabus 2015). (Teaching unit Compulsory)
ECTS credits: 6
Teaching languages: Catalan, Spanish

Teaching staff
Coordinator: Definit a la infoweb de l'assignatura.
Others: Definit a la infoweb de l'assignatura.

Prior skills
Upper secondary school chemistry.

Degree competences to which the subject contributes

Specific:
1. CE 4 AERO. Capacidad para comprender y aplicar los principios de conocimientos básicos de la química general, química orgánica e inorgánica y sus aplicaciones en la ingeniería. (CIN/308/2009, BOE 18.2.2009)

General:
4. EFFICIENT USE OF EQUIPMENT AND INSTRUMENTS - Level 1: Using instruments, equipment and software from the laboratories of general or basic use. Realising experiments and proposed practices and analyzing obtained results.

Transversal:
2. SUSTAINABILITY AND SOCIAL COMMITMENT - Level 1. Analyzing the world’s situation critically and systemically, while taking an interdisciplinary approach to sustainability and adhering to the principles of sustainable human development. Recognizing the social and environmental implications of a particular professional activity.
3. TEAMWORK - Level 1. Working in a team and making positive contributions once the aims and group and individual responsibilities have been defined. Reaching joint decisions on the strategy to be followed.

Teaching methodology
Directed learning in large groups consists of lectures in which the lecturer will present the main content of the subject and encourage students to participate in exercises and practical case studies, to ensure that all students take an active role in the learning process and consolidate their understanding of the theoretical material.
Directed learning in small groups consists of practical laboratory sessions, during which students will generally work in pairs. Practical sessions are designed to consolidate the theoretical concepts studied in lectures and to develop basic laboratory skills, as well as strengthening team work skills, a generic competence acquired during the course.
After most sessions, students will be set work to complete in their own time, which may include recommended reading or questions and problem-solving exercises to be completed individually or in groups, which form the basis of the self-directed learning for this subject.
On completion of Chemistry, students will be able to:
- Identify, formulate and solve problems that relate electronic and atomic structure to the chemical and physical properties of elements and different types of bonds. Identify the different crystal networks of solid materials. Define the possible imperfections of solid materials. Explain solid-solid diffusion and its dependence on temperature. Identify the industrial applications of solid-solid diffusion.
- Explain the process of corrosion and identify potentially corrosive materials on the basis of their chemical properties. Complete exercises and solve problems on corrosion and understand its importance in the aeronautics industry.
- Explain the process of combustion and identify potentially combustible substances on the basis of their chemical properties. Complete exercises and solve problems involving chemical kinetics and combustion models.

### Study load

<table>
<thead>
<tr>
<th>Total learning time: 150h</th>
<th>Hours large group:</th>
<th>26h</th>
<th>17.33%</th>
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<tbody>
<tr>
<td></td>
<td>Hours medium group:</td>
<td>13h</td>
<td>8.67%</td>
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<tr>
<td></td>
<td>Hours small group:</td>
<td>13h</td>
<td>8.67%</td>
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<tr>
<td></td>
<td>Guided activities:</td>
<td>14h</td>
<td>9.33%</td>
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<tr>
<td></td>
<td>Self study:</td>
<td>84h</td>
<td>56.00%</td>
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</table>
# Content

## Atomic Structure and Bonding

**Learning time:** 25h  
Theory classes: 6h  
Practical classes: 3h  
Laboratory classes: 0h  
Guided activities: 2h  
Self study: 14h

**Description:**
1.1 Atomic numbers, electronic and atomic structure  
1.2 Periodic properties  
1.3 Types of bonds  
1.4 Nuclear disintegration phenomena

**Related activities:**
Activity 1. Individual assessment tests. Test 1  
Activity 2. Practical computer work  
Activity 3. Directed Activities 5 and 6

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## Crystal Structure, Imperfections in Solid Materials and Diffusion

**Learning time:** 42h  
Theory classes: 10h  
Practical classes: 5h  
Laboratory classes: 3h  
Guided activities: 4h  
Self study: 20h

**Description:**
2.1 Unit cell  
2.2 Crystal systems and structures  
2.3 Crystallographic directions and planes  
2.4 Point defects  
2.5 Imperfections  
2.6 The diffusion phenomenon and mechanisms; Stationary and nonstationary diffusion  
2.7 Temperature dependence of diffusion  
2.8 Solid solutions; Phase diagrams; Alloys

**Related activities:**
Activity 1. Individual assessment tests. Test 1  
Activity 2. Laboratory activities (computer room)  
Activity 3. Directed Activities 7, 8 and 9
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**Electrochemistry**

**Learning time:** 42h

- Theory classes: 5h
- Practical classes: 3h
- Laboratory classes: 5h
- Guided activities: 4h
- Self study: 25h

**Description:**
3.1 Redox reactions
3.2 Cell potential
3.3 Free energy and electrical work
3.4 Electrochemical process in aerospace fuel cells
3.5 Corrosion: Corrosive agents and antioxidants

**Related activities:**
Activity 1. Individual assessment tests. Test 2
Activity 2. Laboratory activities
Activity 3. Directed Activities 10 and 11

**Combustion**

**Learning time:** 41h

- Theory classes: 5h
- Practical classes: 2h
- Laboratory classes: 5h
- Guided activities: 4h
- Self study: 25h

**Description:**
4.1 Reaction mechanisms
4.2 Chemical kinetics
4.3 Chain reactions; Explosions
4.4 Temperature dependence of reaction rate
4.5 Combustion: Examples of combustion in aircraft motors

**Related activities:**
Activity 1. Individual assessment tests. Test 2
Activity 2. Laboratory activities
Activity 3. Directed Activities 12 and 13
### Planning of activities

<table>
<thead>
<tr>
<th>INDIVIDUAL ASSESSMENT TEST. TEST 1</th>
<th>Hours: 1h</th>
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<tbody>
<tr>
<td></td>
<td>Theory classes: 1h</td>
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<tr>
<td></td>
<td>Practical classes: 0h</td>
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<tr>
<td></td>
<td>Laboratory classes: 0h</td>
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<td></td>
<td>Guided activities: 0h</td>
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<td></td>
<td>Self study: 0h</td>
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**Description:**
Short-answer questions and discursive questions on the theory content covered in class, laboratory sessions and computer work. Students will be tested on the topics Atomic Structure and Bonding and Crystal Structure, Imperfections in Solid Materials and Diffusion.

**Support materials:**
Tables and calculator.

**Descriptions of the assignments due and their relation to the assessment:**
Completion of the test by each student.
Lecturers' assessment of the directed and self-directed learning outcomes for each student.
The results will count towards the final mark.

**Specific objectives:**
The test will focus on students' knowledge of the following points:

a) The periodic properties of the elements and their associated macroscopic properties.
b) The characteristics of different types of bonds and their associated macroscopic properties.
c) The parameters for defining a crystallographic structure and the different crystal networks.
d) The characteristics of diffusion mechanisms and their industrial applications.
e) The characteristics of different types of imperfections in solid materials and their industrial applications and effects.

<table>
<thead>
<tr>
<th>INDIVIDUAL ASSESSMENT TEST. TEST 2</th>
<th>Hours: 1h</th>
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<tbody>
<tr>
<td></td>
<td>Theory classes: 1h</td>
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<tr>
<td></td>
<td>Practical classes: 0h</td>
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<td></td>
<td>Laboratory classes: 0h</td>
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<td>Guided activities: 0h</td>
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<td>Self study: 0h</td>
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</tbody>
</table>

**Description:**
Short-answer questions and discursive questions on the theory content covered in class, laboratory sessions and computer work. Students will be tested on the topics Electrochemistry and Combustion.

**Support materials:**
Short-answer questions and discursive questions on the theory content covered in class, laboratory sessions and computer work. Students will be tested on the topics Electrochemistry and Combustion.

**Descriptions of the assignments due and their relation to the assessment:**
Completion of the test by each student.
Lecturers' assessment of the directed and self-directed learning outcomes for each student.
The results will count towards the final mark.
LABORATORY

Description:
Two-hour practical laboratory sessions. Attendance of all laboratory sessions is compulsory. Students will work in pairs.
Laboratory work focuses on the practical application of the material covered in Electrochemistry and Combustion.

Support materials:
Required materials and reagents.
Detailed description of each activity and accompanying questions.

Descriptions of the assignments due and their relation to the assessment:
Lecturers' assessment of the directed learning outcomes for each student.
The results will count towards the overall mark for practical activities. Students must submit a report on each practical activity for assessment. The combined results will count for 10% of the final mark.

Specific objectives:
At the end of the series of practical sessions, students will be able to:
- carry out laboratory work in compliance with environmental and safety guidelines;
- understand the importance of organisation in laboratory work;
- handle laboratory material and chemical samples correctly;
- use laboratory equipment correctly;
- assess the value of their own results and compare them with the work of other groups;
- identify free energy and electrical work in an electrochemical process in practical cases;
- recognise the most corrosive agents and the main antioxidant elements and/or products;
- carry out redox reactions and monitor the main parameters;
- use controlled laboratory reactions to infer the temperature dependence of the rate of chemical reactions;
- carry out controlled combustions and monitor the main parameters.

PRACTICAL COMPUTER WORK

Hours: 6h
Theory classes: 0h
Practical classes: 0h
Laboratory classes: 3h
Guided activities: 0h
Self study: 3h
ACTIVITY 5: PERIODIC PROPERTIES OF CHEMICAL ELEMENTS

**Description:**
Groups of 10 students will infer the period properties of different elements on the basis of their positions in the periodic table. This activity will be carried out in a cooperative learning room.

**Support materials:**
Periodic table of the elements.

**Descriptions of the assignments due and their relation to the assessment:**
Students will present reasoned explanations of the relationships between the periodic properties of the elements, the position of each element in the periodic table, and the macroscopic characteristics associated with these properties. Proportion of final mark: 2%.

**Specific objectives:**
Establish the relationship between the periodic properties of an element and its electronic configuration.

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**ACTIVITY 6: CHEMICAL BONDING**

**Description:**
Groups of 10 students will discuss the potential chemical bonds and interactions of different elements on the basis of their position in the periodic table. This activity will be carried out in a cooperative learning room.

**Support materials:**
Periodic table of the elements. Three-dimensional molecular models.
ACTIVITY 8: SOLID-SOLID DIFFUSION

Description:
Groups of 10 students will complete and discuss exercises and specific problems involving solid-solid diffusion and its dependence on temperature. This activity will be carried out in a cooperative learning room.

Support materials:
Descriptions of the exercises and problems in digital and paper format.

Specific objectives:
Predict the types of bonds and chemical reactions an element can form on the basis of its periodic properties.

ACTIVITY 9: DIFFUSION IN SOLIDS

Hours: 2h
Guided activities: 2h

Description:
Groups of 10 students will construct three-dimensional models of the most representative unit cells. This activity will be carried out in a cooperative learning room.

Support materials:
Three-dimensional molecular models.

Specific objectives:
Predict the types of bonds and chemical reactions an element can form on the basis of its periodic properties.
**ACTIVITY 11: PRACTICAL STUDY OF CORROSION PHENOMENA**

**Description:**
Groups of 20 students will discuss and assess the results of practical sessions on redox experiments. Laboratory results will be used as the basis for discussion of the effects of corrosion and the most important industrial effects of this phenomenon. This activity will be carried out in a cooperative learning room.

**Support materials:**
Results of the practical sessions carried out by all groups.

**Specific objectives:**
Students will compare their results with the work of other groups, as well as submitting their data to statistical processing. Guidance will be provided on how to write the reports on practical sessions.

**Hours:** 2h
Guided activities: 2h

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**ACTIVITY 12: COMBUSTION AND EXPLOSION**

**Description:**
Groups of 10 students will complete and discuss exercises and specific problems involving different combustion, explosion and chemical kinetic reactions. This activity will be carried out in a cooperative learning room.

**Support materials:**
Descriptions of the exercises and problems in digital and paper format.

**Descriptions of the assignments due and their relation to the assessment:**
Students will present their solutions to the exercises and problems. Proportion of final mark: 2%.

**Specific objectives:**
The teaching staff will assess students' understanding of combustion reactions and explosion mechanisms and their industrial applications in the aerospace sector.

**Hours:** 2h
Guided activities: 2h

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**ACTIVITY 13: PRACTICAL STUDY OF COMBUSTION**

**Description:**
Groups of 20 students will discuss and assess the results of practical sessions on combustion and chemical kinetics experiments. Laboratory results will be used as the basis for discussion of the main industrial applications. This activity will be carried out in a cooperative learning room.

**Support materials:**
Results of the practical sessions carried out by all groups.

**Hours:** 2h
Guided activities: 2h

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**ACTIVITY 10: REDOX REACTIONS**

**Description:**
Groups of 10 students will complete and discuss exercises and specific problems involving different redox reactions.
This activity will be carried out in a cooperative learning room.

**Support materials:**
Descriptions of the exercises and problems in digital and paper format.

**Descriptions of the assignments due and their relation to the assessment:**
Students will present their solutions to the exercises and problems. Proportion of final mark: 2%.

**Specific objectives:**
The teaching staff will assess students' understanding and practical handling of redox reaction mechanisms.

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**Specific objectives:**
Students will compare their results with the work of other groups, as well as submitting their data to statistical processing. Guidance will be provided on how to write the reports on practical sessions.

**ACTIVITY 10: REDOX REACTIONS**

**Hours:** 2h
Guided activities: 2h

**Qualification system**
The final mark for this subject will be calculated from:
Tests: 20% (10% each)
Mid-semester examination: 20%
Final examination: 30%
Laboratory sessions: 20%
Directed activities: 10%

**Regulations for carrying out activities**
Students must bring the description of practical activities to all laboratory sessions and comply with health and safety regulations. Attendance of all laboratory sessions is compulsory. Failure to attend one or more sessions will result in the award of a mark of 0 (zero) for the corresponding report.
Students will carry out all practical laboratory work in pairs. Practical computer work must be completed individually.
Bibliography

Basic:


Complementary:
