Course guides  
390105 - FQ1 - Chemistry I

Unit in charge: Barcelona School of Agri-Food and Biosystems Engineering
Teaching unit: 745 - DEAB - Department of Agri-Food Engineering and Biotechnology.

Degree: 
BACHELOR’S DEGREE IN AGRICULTURAL ENGINEERING (Syllabus 2009). (Compulsory subject).
BACHELOR’S DEGREE IN AGRICULTURAL, ENVIRONMENTAL AND LANDSCAPE ENGINEERING (Syllabus 2009). (Compulsory subject).
BACHELOR’S DEGREE IN FOOD ENGINEERING (Syllabus 2009). (Compulsory subject).
BACHELOR’S DEGREE IN BIOSYSTEMS ENGINEERING (Syllabus 2009). (Compulsory subject).
BACHELOR’S DEGREE IN AGRONOMIC SCIENCE ENGINEERING (Syllabus 2018). (Compulsory subject).

Academic year: 2020  ECTS Credits: 6.0  Languages: Catalan

LECTURER
Coordinating lecturer: PATRICIA JIMENEZ DE RIDDER
Others: TERESA BALANYÀ MARTÍ - JOSE SABATE REBOLL - PATRICIA JIMENEZ DE RIDDER - MARIA TERESA COLL AUSIO

DEGREE COMPETENCES TO WHICH THE SUBJECT CONTRIBUTES
Specific:  
1. Knowledge of the basic concepts of general chemistry, inorganic and organic chemistry, and ability to use them in engineering applications.

TEACHING METHODOLOGY
The directed learning time with a large group consist, first, in lectures where teachers do a brief explanation to introduce learning objectives related to the basic concepts of the subject. Later teacher motivate and engage students to actively participate in their learning through exercises. On the other hand, hours of directed learning in small groups consist of making classes solving problems and numerical exercises or labs. The practices are designed as application of theoretical concepts in the laboratory. In general, before and after each session tasks outside the classroom are proposed, such as resolution of issues and problems that are the basis of learning and self-guided.

LEARNING OBJECTIVES OF THE SUBJECT
At the end of the Chemistry 1 course, the student should be able to:
- Determine the spontaneity of a chemical reaction from the concepts of chemical thermodynamics.
- Explain the meaning of chemical equilibrium and the effect they may exert different variables (concentrations of reactants and products, total pressure and temperature). Relate the equilibrium constant with the standard Gibbs free energy.
- Identifying and solving the necessary equations to perform calculations of balance: balance of mass and charge, and equilibrium constants.
- Identify requirements for volumetric analysis, interpreted the valuation curves and make the prior and end calculations.
- Perform correctly basic measures of a chemistry lab.
STUDY LOAD

<table>
<thead>
<tr>
<th>Type</th>
<th>Hours</th>
<th>Percentage</th>
</tr>
</thead>
<tbody>
<tr>
<td>Hours small group</td>
<td>20.0</td>
<td>13.33</td>
</tr>
<tr>
<td>Hours large group</td>
<td>40.0</td>
<td>26.67</td>
</tr>
<tr>
<td>Self study</td>
<td>90.0</td>
<td>60.00</td>
</tr>
</tbody>
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**Total learning time:** 150 h

CONTENTS

**BASIC CONCEPTS AND IONIC SOLUTIONS**

**Description:**
1.1. Types of chemical compounds and inorganic compounds formulation
1.2. Electrolytes strong solutions and calculations of the ionic concentrations

**Related activities:**
Activity 1: Class of theoretical explanation.
Activity 2: Individual final assessment test.
Activity 3: Laboratory activities.

**Full-or-part-time:** 28h
Theory classes: 7h
Laboratory classes: 6h
Self study: 15h

**THERMODYNAMICS AND CHEMICAL EQUILIBRIUM**

**Description:**
3.1. Energy and spontaneity of chemical reactions.
3.2. Equilibrium constant.
3.3. Modification of the equilibrium conditions
3.4 Oxidation reduction reactions

**Related activities:**
Activity 1: Class of theoretical explanation
Activity 2: Individual final assessment test

**Full-or-part-time:** 46h
Theory classes: 15h
Laboratory classes: 6h
Self study: 25h
BALANCE IN AQUEOUS MEDIUM

**Description:**
4.1 Acid-base reactions  
4.2 Precipitation and complexation reactions  
4.3 Equilibrium reactions in aqueous medium  
4.4 Volumetric analysis

**Related activities:**
Activity 1: Class of theoretical explanation.  
Activity 2: Individual final assessment test  
Activity 3: Lab activities.

**Full-or-part-time:** 61h  
Theory classes: 18h  
Laboratory classes: 8h  
Self study: 35h

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**ACTIVITIES**

**(ENG) ACTIVITY 1: THEORY LESSONS**

**Full-or-part-time:** 36h  
Theory classes: 36h

**(ENG) ACTIVITY 2: INDIVIDUAL TESTS**

**Full-or-part-time:** 4h  
Theory classes: 4h

**(ENG) ACTIVITY 3: LABORATORY**

**Full-or-part-time:** 50h  
Laboratory classes: 20h  
Self study: 30h
GRADING SYSTEM

It will be made two individual tests in the classroom: a first test (P1) will include the first half of the subject matter and a second test that will have two modes, the second half test (P2) or final test (F).

The second half test is focused in the second half of the subject matter, though occasionally it may be necessary to use any concept or procedure of the first half. The final test includes all the subject matter. All test also include all the subject matter of practice.

Students, who have obtained a P1 test rating equal or greater than 4, may choose to make second or final test. All other students will have the final test. The final grade for the course, Nfinal, is obtained with one of the following ways:

N1: qualification of the P1 test.
N2: qualification of the second test P2
N3: qualification of the final test
N4: Qualification of small group activities
CG: Qualification of generic competition.

N_{\text{final}} = 0.30N_1 + 0.45N_2 + 0.20N_4 + 0.05CG
or:
N_{\text{final}} = 0.75N_3 + 0.20N_4 + 0.05CG

In the case of failing the course with a final course grade greater than NP, the final test (F) may be re-evaluated.

BIBLIOGRAPHY

Basic: